

## 1

**The Statistical Nature of Radiation, Emission, and Interaction**

## 1.1

**Introduction and Scope**

This book is about statistics, with emphasis on its role in radiation physics, measurements, and radiation protection. That this subject is essential for understanding in these areas stems directly from the statistical nature of the submicroscopic, atomic world, as we briefly discuss in the next section. The principal aspects of atomic physics with which we shall be concerned are radioactive decay, radiation transport, and radiation interaction. Knowledge of these phenomena is necessary for success in many practical applications, which include dose assessment, shielding design, and the interpretation of instrument readings. Statistical topics will be further developed for establishing criteria to measure and characterize radioactive decay, assigning confidence limits for measured quantities, and formulating statistical measures of performance and compliance with regulations. An introduction to biological dose–response relations and to modeling the biological effects of radiation will also be included.

## 1.2

**Classical and Modern Physics – Determinism and Probabilities**

A principal objective of physical science is to discover laws and regularities that provide a quantitative description of nature as verified by observation. A desirable and useful outcome to be derived from such laws is the ability to make valid predictions of future conditions from a knowledge of the present state of a system. Newton's classical laws of motion, for example, determine completely the future motion of a system of objects if their positions and velocities at some instant of time and the forces acting between them are known. On the scale of the very large, the motion of the planets and moons can thus be calculated forward (and backward) in time, so that eclipses and other astronomical phenomena can be predicted with great accuracy. On the scale of everyday common life, Newton's laws describe all manner of diverse experience involving motion and statics. However, in the early twentieth century, the seemingly inviolate tenets of traditional physics were found to fail on the small scale

of atoms. In place of a deterministic world of classical physics, it was discovered that atoms and radiation are governed by definite, but statistical, laws of quantum physics. Given the present state of an atomic system, one can predict its future, but only in statistical terms. What is the probability that a radioactive sample will undergo a certain number of disintegrations in the next minute? What is the probability that a given 100-keV gamma photon will penetrate a 0.5-cm layer of soft tissue? According to modern quantum theory, these questions can be answered as fully as possible only by giving the complete set of *probabilities* for obtaining any possible result of a measurement or observation.

By the close of the nineteenth century, the classical laws of mechanics, electromagnetism, thermodynamics, and gravitation were firmly established in physics. There were, however, some outstanding problems – evidence that all was not quite right. Two examples illustrate the growing difficulties. First, in the so-called “ultraviolet catastrophe,” classical physics incorrectly predicted the distribution of wavelengths in the spectrum of electromagnetic radiation emitted from hot bodies, such as the sun. Second, sensitive measurements of the relative speed of light in different directions on earth – expected to reveal the magnitude of the velocity of the earth through space – gave a null result (no difference!). Planck found that the first problem could be resolved by proposing a nonclassical, quantum hypothesis related to the emission and absorption of radiation by matter. The now famous quantum of action,  $h = 6.6261 \times 10^{-34}$  J s, was thus introduced into physics. The second dilemma was resolved by Einstein in 1905 with the revolutionary special theory of relativity. He postulated that the speed of light has the same numerical value for all observers in uniform translational motion with respect to one another, a situation wholly in conflict with velocity addition in Newtonian mechanics. Special relativity further predicts that energy and mass are equivalent and that the speed of light in a vacuum is the upper limit for the speed that any object can have. The classical concepts of absolute space and absolute time, which had been accepted as axiomatic tenets for Newton’s laws of motion, were found to be untenable experimentally.

### ■ Example

In a certain experiment, 1000 monoenergetic photons are normally incident on a shield. Exactly 276 photons are observed to interact in the shield, while 724 photons pass through without interacting.

- a) What is the probability that the next incident photon, under the same conditions, will not interact in the shield?
- b) What is the probability that the next photon *will* interact?

### Solution

- a) Based on the given data, we estimate that the probability for a given photon to traverse the shield with no interaction is equal to the observed fraction that did not interact. Thus, the “best value” for the probability  $\text{Pr}(\text{no})$  that the next photon will pass through without interacting is

$$\text{Pr}(\text{no}) = \frac{724}{1000} = 0.724. \quad (1.1)$$

- b) By the same token, the estimated probability  $\text{Pr}(\text{yes})$  that the next photon will interact is

$$\text{Pr}(\text{yes}) = \frac{276}{1000} = 0.276, \quad (1.2)$$

based on the observation that 276 out of 1000 interacted.

This example suggests several aspects of statistical theory that we shall see often throughout this book. The sum of the probabilities for all possible outcomes considered in an experiment must add up to unity. Since only two possible alternatives were regarded in the example – either a photon interacted in the shield or it did not – we had  $\text{Pr}(\text{no}) + \text{Pr}(\text{yes}) = 1$ . We might have considered further whether an interaction was photoelectric absorption, Compton scattering, or pair production. We could assign separate probabilities for these processes and then ask, for example, what the probability is for the next interacting photon to be Compton scattered in the shield. In general, whatever number and variety of possible outcomes we wish to consider, the sum of their probabilities must be unity. This condition thus requires that there be *some* outcome for each incident photon.

It is evident, too, that a larger data sample will generally enable more reliable statistical predictions to be made. Knowing the fate of 1000 photons in the example gives more confidence in assigning values to the probabilities  $\text{Pr}(\text{no})$  and  $\text{Pr}(\text{yes})$  than would knowing the fate of, say, only 10 photons. Having data for  $10^8$  photons would be even more informative.

Indeed, the general question arises, “How can one ever know the actual, true numerical values for many of the statistical quantities that we must deal with?” Using appropriate samples and protocols that we shall develop later, one can often obtain rather precise values, but always within well-defined statistical limitations. A typical result expresses a “best” numerical value that lies within a given range with a specified degree of confidence. For instance, from the data given in the example above, we can express the “measured” probability of no interaction as

$$\text{Pr}(\text{no}) = 0.724 \pm 0.053 \quad (95\% \text{ confidence level}). \quad (1.3)$$

(The stated uncertainty,  $\pm 0.053$ , is  $\pm 1.96$  standard deviations from an estimated mean of 0.724, based on the single experiment, as we shall discuss later in connection with the normal distribution.) Given the result (1.3), there is still no guarantee that the “true” value is actually in the stated range. Many such probabilities can also be accurately calculated from first principles by using quantum mechanics. In all known instances, the theoretical results are in agreement with measurements. Confirmation by observation is, of course, the final criterion for establishing the validity of the properties we ascribe to nature.

### 1.3

#### Semiclassical Atomic Theory

Following the unexpected discovery of X-rays by Roentgen in 1895, a whole series of new findings ushered in the rapidly developing field of atomic and radiation

physics. Over the span of the next two decades, it became increasingly clear that classical science did not give a correct picture of the world as new physics unfolded. Becquerel discovered radioactivity in 1896, and Thomson measured the charge-to-mass ratio of the electron in 1897. Millikan succeeded in precisely measuring the charge of the electron in 1909. By 1910, a number of radioactive materials had been investigated, and the existence of isotopes and the transmutation of elements by radioactive decay were recognized. In 1911, Rutherford discovered the atomic nucleus – a small, massive dot at the center of the atom, containing all of the positive charge of the neutral atom and virtually all of its mass. The interpretation of his experiments on alpha-particle scattering from thin layers of gold pointed to a planetary structure for an atom, akin to a miniature solar system. The atom was pictured as consisting of a number of negatively charged electrons traversing most of its volume in rapid orbital motion about a tiny, massive, positively charged nucleus.

The advance made with the discovery of the nuclear atom posed another quandary for classical physics. The same successful classical theory (Maxwell's equations) that predicted many phenomena, including the existence of electromagnetic radiation, required the emission of energy by an accelerated electric charge. An electron in orbit about a nucleus should thus radiate energy and quickly spiral into the nucleus. The nuclear atom could not be stable. To circumvent this dilemma, Bohr in 1913 proposed a new, semiclassical nuclear model for the hydrogen atom. The single electron in this system moved in classical fashion about the nucleus (a proton). However, in nonclassical fashion Bohr postulated that the electron could occupy only certain circular orbits in which its angular momentum about the nucleus was quantized. (The quantum condition specified that the angular momentum was an integral multiple of Planck's constant divided by  $2\pi$ .) In place of the continuum of unstable orbits allowed by classical mechanics, the possible orbits for the electron in Bohr's model were discrete. Bohr further postulated that the electron emitted radiation only when it went from one orbit to another of lower energy, closer to the nucleus. The radiation was then emitted in the form of a photon, having an energy equal to the difference in the energy the electron had in the two orbits. The atom could absorb a photon of the same energy when the electron made the reverse transition between orbits. These criteria for the emission and absorption of atomic radiation replaced the classical ideas. They also implied the recognized fact that the chemical elements emit and absorb radiation at the same wavelengths and that different elements would have their own individual, discrete, characteristic spectra. Bohr's theory for the hydrogen atom accounted in essential detail for the observed optical spectrum of this element. When applied to other atomic systems, however, the extension of Bohr's ideas often led to incorrect results.

An intensive period of semiclassical physics then followed into the 1920s. The structure and motion of atomic systems was first described by the equations of motion of classical physics, and then quantum conditions were superimposed, as Bohr had done for hydrogen. The quantized character of many variables, such as energy and angular momentum, previously assumed to be continuous, became increasingly evident experimentally.

Furthermore, nature showed a puzzling wave–particle duality in its fundamental makeup. Electromagnetic radiation, originally conceived as a purely wave phenomenon, exhibited properties of both waves and particles. The diffraction and interference of X-rays was demonstrated experimentally by von Laue in 1912, establishing their wave character. Einstein’s explanation of the photoelectric effect in 1905 described electromagnetic radiation of frequency  $\nu$  as consisting of packets, or photons, having energy  $E = h\nu$ . The massless photon carries an amount of momentum that is given by the relation

$$p = \frac{E}{c} = \frac{h\nu}{c}, \quad (1.4)$$

where  $c = 2.9979 \times 10^8 \text{ m s}^{-1}$  is the speed of light in a vacuum. This particle-like property of momentum is exhibited experimentally, for example, by the Compton scattering of photons from electrons (1922). The wavelength  $\lambda$  of the radiation is given by  $\lambda = c/\nu$ . It follows from Eq. (1.4) that the relationship between the wavelength and momentum of a photon is given by

$$\lambda = \frac{h}{p}. \quad (1.5)$$

In 1924, de Broglie postulated that this relationship applies not only to photons, but also to other fundamental atomic particles. Electron diffraction was demonstrated experimentally by Davisson and Germer in 1927, with the electron wavelength being correctly given by Eq. (1.5). (Electron microscopes have much shorter wavelengths and hence much greater resolving power than their optical counterparts.)

There was no classical analogue to these revolutionary quantization rules and the wave–particle duality thus introduced into physics. Yet they appeared to work. The semiclassical procedures had some notable successes, but they also led to some unequivocally wrong predictions for other systems. There seemed to be elements of truth in quantizing atomic properties, but nature’s secrets remained hidden in the early 1920s.

## 1.4

### Quantum Mechanics and the Uncertainty Principle

Heisenberg reasoned that the root of the difficulties might lie in the use of nonobservable quantities to describe atomic constituents – attributes that the constituents might not even possess. Only those properties should be ascribed to an object that have an operational definition through an experiment that can be carried out to observe or measure them. What does it mean, for example, to ask whether an electron is blue or red, or even to ask whether an electron has a color? Such questions must be capable of being answered by experiment, at least in principle, or else they have no meaning in physics. Using only observable atomic quantities, such as those associated with the frequencies of the radiation emitted by an atom, Heisenberg in 1924 developed a new, matrix theory of quantum mechanics. At almost the same time, Schrödinger formulated his wave equation from an entirely

different standpoint. He soon was able to show that his formulation and Heisenberg's were completely equivalent. The new quantum mechanics was born.

In the Newtonian mechanics employed by Bohr and others in the semiclassical theories, it was assumed that an atomic electron possesses a definite position and velocity at every instant of time. Heisenberg's reasoning required that, in order to have any meaning or validity, the very concept of the "position and velocity of the electron" should be defined operationally by means of an experiment that would determine it. He showed that the act of measuring the position of an electron ever more precisely would, in fact, make the simultaneous determination of its momentum (and hence velocity) more and more uncertain. In principle, the position of an electron could be observed experimentally by scattering a photon from it. The measured position would then be localized to within a distance comparable to the wavelength of the photon used, which limits its spatial resolving power. The scattered photon would, in turn, impart momentum to the electron being observed. Because of the finite aperture of any apparatus used to detect the scattered photon, its direction of scatter and hence its effect on the struck electron's momentum would not be known exactly. To measure the position of the electron precisely, one would need to use photons of very short wavelength. These, however, would have large energy and momentum, and the act of scattering would be coupled with large uncertainty in the simultaneous knowledge of the electron's momentum. Heisenberg showed that the product of the uncertainties in the position  $\Delta x$  in any direction in space and the component of momentum  $\Delta p_x$  in that direction must be at least as large as Planck's constant divided by  $2\pi$  ( $\hbar = h/2\pi = 1.0546 \times 10^{-34}$  J s):

$$\Delta x \Delta p_x \geq \hbar. \quad (1.6)$$

It is thus impossible to assign both position and momentum simultaneously with unlimited precision. (The equality applies only under optimum conditions.) The inequality (1.6) expresses one form of Heisenberg's uncertainty principle. A similar relationship exists between certain other pairs of variables, such as energy  $E$  and time  $t$ :

$$\Delta E \Delta t \geq \hbar. \quad (1.7)$$

The energy of a system cannot be determined with unlimited precision within a short interval of time.

These limits imposed by the uncertainty principle are not due to any shortcomings in our measurement techniques. They simply reflect the way in which the act of observation itself limits simultaneous knowledge of certain pairs of variables. To speculate whether an electron "really does have" an exact position and velocity at every instant of time, although we cannot know them together, apparently has no operational meaning. As we shall see in an example below, the limits have no practical effect on massive objects, such as those experienced in everyday life. In contrast, however, on the atomic scale the limits reflect an essential need to define carefully and operationally the concepts that are to have meaning and validity.

The subsequent development of quantum mechanics has provided an enormously successful quantitative description of many phenomena: atomic and nuclear struc-

ture, radioactive decay, lasers, semiconductors, antimatter, electron diffraction, superconductivity, elementary particles, radiation emission and absorption, the covalent chemical bond, and many others. It has revealed the dual wave–particle nature of the constituents of matter. Photons, electrons, neutrons, protons, and other particles have characteristics of both particles and waves. Instead of having a definite position and velocity, they can be thought of as being “smeared out” in space, as reflected by the uncertainty principle. They can be described in quantum mechanics by wave packets related to a probability density for observing them in different locations. They have both momentum  $p$  and wavelength  $\lambda$ , which are connected by the de Broglie relation (1.5). Endowed with such characteristics, the particles exhibit diffraction and interference effects under proper experimental conditions. Many quantum-mechanical properties, essential for understanding atomic and radiation physics, simply have no classical analogue in the experience of everyday life.

#### Example

The electron in the hydrogen atom is localized to within about  $1 \text{ \AA}$ , which is the size of the atom. Use the equality in the uncertainty relation to estimate the uncertainty in its momentum. Estimate the order of magnitude of the kinetic energy that the electron (mass  $= m = 9.1094 \times 10^{-31} \text{ kg}$ ) would have in keeping with this amount of localization in its position.

#### Solution

With  $\Delta x = 1 \text{ \AA} = 10^{-10} \text{ m}$  in Eq. (1.6), we estimate that the uncertainty in the electron’s momentum is<sup>1)</sup>

$$\Delta p \cong \frac{\hbar}{\Delta x} = \frac{1.05 \times 10^{-34} \text{ J s}}{10^{-10} \text{ m}} \cong 10^{-24} \text{ kg m s}^{-1}. \quad (1.8)$$

We assume that the electron’s momentum is about the same order of magnitude as this uncertainty. Denoting the electron mass by  $m$ , we estimate for its kinetic energy

$$T = \frac{(\Delta p)^2}{2m} \cong \frac{(10^{-24} \text{ kg m s}^{-1})^2}{2 \times 9.11 \times 10^{-31} \text{ kg}} \cong 5 \times 10^{-19} \text{ J} \cong 3 \text{ eV}, \quad (1.9)$$

since  $1 \text{ eV} = 1.60 \times 10^{-19} \text{ J}$ . An electron confined to the dimensions of the hydrogen atom would be expected to have a kinetic energy in the eV range. The mean kinetic energy of the electron in the ground state of the hydrogen atom is  $13.6 \text{ eV}$ .

The uncertainty principle requires that observing the position of a particle with increased precision entails increased uncertainty in the simultaneous knowledge of its

1) Energy, which has the dimensions of force  $\times$  distance, has units  $1 \text{ J} = 1 \text{ N m}$ . The newton of force has the same units as mass  $\times$  acceleration:  $1 \text{ N} = 1 \text{ kg m s}^{-2}$ . Therefore,  $1 \text{ J s m}^{-1} = 1 \text{ kg m s}^{-1}$ , which are the units of momentum (mass  $\times$  velocity).

momentum, or energy. Greater localization of a particle, therefore, is accompanied by greater likelihood that measurement of its energy will yield a large value. Conversely, if the energy is known with precision, then the particle must be “spread out” in space. Particles and photons can be described mathematically by quantum-mechanical wave packets, which, in place of classical determinism, provide probability distributions for the possible results of any measurement. These essential features of atomic physics are not manifested on the scale of familiar, everyday objects.

### ■ Example

How would the answers to the last example be affected if

- a) the electron were localized to nuclear dimensions ( $\Delta x \sim 10^{-15}$  m) or
- b) the electron mass were 100 g?

### Solution

- a) With  $\Delta x \sim 10^{-15}$  m, the equality in the uncertainty principle (1.6) gives, in place of Eq. (1.8),  $\Delta p \cong 10^{-19}$  kg m s<sup>-1</sup>, five orders of magnitude larger than before. The corresponding electron energy would be relativistic. Calculation shows that the energy of an electron localized to within  $10^{-15}$  m would be about 200 MeV. (The numerical solution is given in Section 2.5 in Turner (2007), listed in the Bibliography at the end of this book.)
- b) Since  $\Delta x$  is the same as before ( $10^{-10}$  m),  $\Delta p$  in Eq. (1.8) is unchanged. With  $m = 100$  g = 0.1 kg, the energy in place of Eq. (1.9) is now smaller by a factor of the mass ratio  $(9.11 \times 10^{-31})/0.1 \cong 10^{-29}$ . For all practical purposes, with the resultant extremely small value of  $T$ , the uncertainty in the velocity is negligible. Whereas the actual electron, localized to such small dimensions, has a large uncertainty in its momentum, the “100-g electron” would appear to be stationary.

Quantum-mechanical effects are generally expressed to a lesser extent with relatively massive objects, as this example shows. By atomic standards, objects in the macroscopic world are massive and have very large momenta. Their de Broglie wavelengths, expressed by Eq. (1.5), are vanishingly small. Quantum mechanics becomes important on the actual physical scale because of the small magnitude of Planck’s constant.

## 1.5

### Quantum Mechanics and Radioactive Decay

Before the discovery of the neutron by Chadwick in 1932, it was speculated that the atomic nucleus must be made up of the then known elementary subatomic particles: protons and electrons. However, according to quantum mechanics, this assumption leads to the wrong prediction of the angular momentum for certain nuclei. The nucleus of  ${}^6\text{Li}$ , for example, would consist of six protons and three electrons, representing nine particles of half-integral spin. By quantum rules for addition of the spins of an odd number of such particles, the resulting nuclear



angular momentum for  ${}^6\text{Li}$  would also have to be a half-integral multiple of Planck's constant,  $\hbar$ . The measured value, however, is just  $\hbar$ . A similar circumstance occurs for  ${}^{14}\text{N}$ . These two nuclei contain *even* numbers (6 and 14) of spin-1/2 particles (protons and neutrons), and hence must have integral angular momentum, as observed.

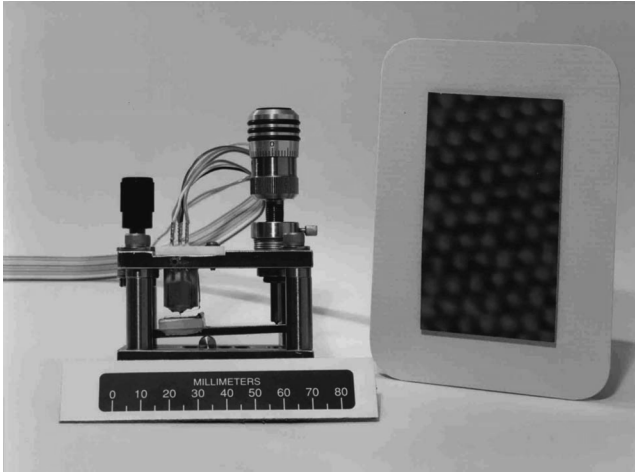
The existence of electrons in the nucleus would also have to be reconciled with the uncertainty principle. In part (a) of the last example, we saw that an electron confined to nuclear dimensions would have an expected kinetic energy in the range of 200 MeV. There is no experimental evidence for such large electron energies associated with beta decay or other nuclear phenomena.

If the electron is not there initially, how is its ejection from the nucleus in beta decay to be accounted for? Quantum mechanics explains the emission of the beta particle through its creation, along with an antineutrino, at the moment of the decay. Both particles are then ejected from the nucleus, causing it to recoil (slightly, because the momentum of the ejected mass is small). The beta particle, the antineutrino, and the recoil nucleus share the released energy, which is equivalent to the loss of mass ( $E = mc^2$ ) that accompanies the radioactive transformation. Since the three participants can share this energy in a continuum of ways, beta particles emitted in radioactive decay have a continuous energy spectrum, which extends out to the total energy released. Similarly, gamma-ray or characteristic X-ray photons are not “present” in the nucleus or atom before emission. They are created when the quantum transition takes place. An alpha particle, on the other hand, is a tightly bound and stable structure of two protons and two neutrons within the nucleus. Alpha decay is treated quantum mechanically as the tunneling of the alpha particle through the confining nuclear barrier, a process that is energetically forbidden in classical mechanics. The emitted alpha particle and recoil nucleus, which separate back to back, share the released energy uniquely in inverse proportion to their masses. The resultant alpha-particle energy spectra are therefore discrete, in contrast to the continuous beta-particle spectra. The phenomenon of tunneling, which is utilized in a number of modern electronic and other applications, is purely quantum mechanical. It has no classical analogue (see Figure 1.1).

The radioactive decay of atoms and the accompanying emission of radiation are thus described in detail by quantum mechanics. As far as is known, radioactive decay occurs spontaneously and randomly, without influence from external factors. The energy thus released derives from the conversion of mass into energy, in accordance with Einstein's celebrated equation,  $E = mc^2$ .

#### ■ Example

Each of 10 identical radioactive atoms is placed in a line of 10 separate counters, having 100% counting efficiency. The question is posed, “Which atom will decay first?” How can the question be answered, and how can the answer be verified?



**Figure 1.1** An early scanning tunneling microscope (left) is used to image the electron clouds of individual carbon atoms on the surface of a highly oriented pyrolytic graphite sample. As a whisker tip just above the surface scans it horizontally in two dimensions, electrons tunnel through a classically forbidden barrier to produce a current through the tip. This current is

extremely sensitive to the separation between the tip and the surface. As the separation tends to change according to the surface contours during the scan, negative feedback keeps it constant by moving a micrometer vertically up or down. These actions are translated by computer into the surface picture shown on the right. (Courtesy of R.J. Warmack.)

#### Solution

Since the atoms are identical and decay is spontaneous, the most one can say is that it is equally likely for *any* of the atoms, 1 through 10, to decay first. The validity of this answer, like any other, is to be decided on an objective basis by suitable experiments or observations. To perform such an experiment, in principle a large number of identical sources of 10 atoms could be prepared and then observed to see how many times the first atom to decay in a source is atom 1, atom 2, and so on. One would find a distribution, giving the relative frequencies for each of the 10 atoms that decays first. Because the atoms are identical, the distribution would be expected to show random fluctuations and become relatively flatter with an increasing number of observations.

#### Example

A source consists of 20 identical radioactive atoms. Each has a 90% chance of decaying within the next 24 h.

- What is the probability that all 20 will decay in 24 h?
- What is the probability that none will decay in 24 h?

#### Solution

- The probability that atom 1 will decay in 24 h is 0.90. The probability that atoms 1 and 2 will both decay in 24 h is  $0.90 \times 0.90 = (0.90)^2 = 0.81$ . That is,

if the experiment is repeated many times, atom 2 is expected to decay in 90% of the cases (also 90%) in which atom 1 also decays. By extension, the probability for all atoms to decay in 24 h is

$$(0.90)^{20} = 0.12. \quad (1.10)$$

- b) Since a given atom must either decay or not decay, the probability for not decaying in 24 h is  $1 - 0.90 = 0.10$ . The probability that none of the 20 atoms will decay is

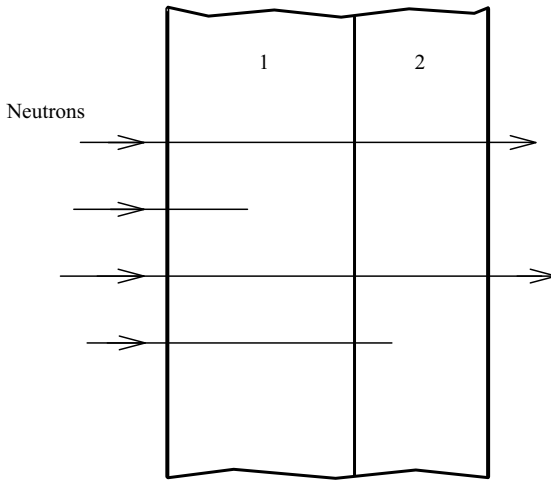
$$(0.10)^{20} = 1.0 \times 10^{-20}. \quad (1.11)$$

As these examples illustrate, quantum mechanics does not generally predict a single, definite result for a single observation. It predicts, instead, a probability for each of all possible outcomes. Quantum mechanics thus brings into physics the idea of the essential randomness of nature. While it is the prevalent conceptual foundation in modern theory, as espoused by Bohr and others, this fundamental role of chance in our universe has not been universally acceptable to all scientists. Which atom will decay first? The contrasting viewpoint was held by Einstein, for example, summed up in the words, “God does not play dice.”

## Problems

- 1.1 The dimensions of angular momentum are those of momentum times distance. Show that Planck's constant,  $h = 6.63 \times 10^{-34}$  J s, has the units of angular momentum.
- 1.2 Einstein's famous equation,  $E = mc^2$ , where  $c$  is the speed of light in a vacuum, gives the energy equivalence  $E$  of mass  $m$ . If  $m$  is expressed in kg and  $c$  in  $\text{m s}^{-1}$ , show that  $E$  is given in J.
- 1.3 According to classical theory, how are electromagnetic waves generated?
- 1.4 Why would the Bohr model of the atom be unstable, according to classical physics?
- 1.5 Calculate the wavelength of a 2.50-eV photon of visible light.
- 1.6 Calculate the wavelength of an electron having an energy of 250 eV.
- 1.7 What is the wavelength of a 1-MeV gamma ray?
- 1.8 If a neutron and an alpha particle have the same speed, how do their wavelengths compare?
- 1.9 If a neutron and an alpha particle have the same wavelength, how do their energies compare?
- 1.10 If a proton and an electron have the same wavelength, how do their momenta compare?
- 1.11 An electron moves freely along the X-axis. According to Eq. (1.6), if the uncertainty in its position in this direction is reduced by a factor of 2, how is the minimum uncertainty in its momentum in this direction affected?

- 1.12 Why is the uncertainty principle, so essential for understanding atomic physics, of no practical consequence for hitting a baseball?
- 1.13 Decay of the nuclide  $^{226}\text{Ra}$  to the ground state of  $^{222}\text{Rn}$  by emission of an alpha particle releases 4.88 MeV of energy.
  - a) What fraction of the total mass available is thus converted into energy? (1 atomic mass unit = 931.49 MeV.)
  - b) What is the initial energy of the ejected alpha particle?
- 1.14 Two conservation laws must be satisfied whenever a radioactive atom decays. As a result of these two conditions, the energies of the alpha particle and the recoil nucleus are uniquely determined in the two-body disintegration by alpha-particle emission. These two laws are also satisfied in beta decay, but do not suffice to determine uniquely the energy of any of the three decay products. What are these two laws, which thus require alpha-particle energies to be discrete and beta-particle energies to be continuous?
- 1.15 The fission of  $^{235}\text{U}$  following capture of a thermal neutron releases an average of 195 MeV. What fraction of the total mass available (neutron plus uranium atom) is thus converted into energy? (1 atomic mass unit = 931.49 MeV.)
- 1.16 Five gamma rays are incident on a concrete slab. Each has a 95% chance of penetrating the slab without experiencing an interaction.
  - a) What is the probability that the first three photons pass through the slab without interacting?
  - b) What is the probability that all five get through without interacting?
- 1.17
  - a) In the last problem, what is the probability that photons 1, 2, and 3 penetrate the slab without interacting, while photons 4 and 5 do not?
  - b) What is the probability that any three of the photons penetrate without interaction, while the other two do not?
- 1.18 Each photon in the last two problems has a binary fate – it either interacts in the slab or else goes through without interaction. A more detailed fate can be considered:  $2/3$  of the photons that interact do so by photoelectric absorption and  $1/3$  that interact do so by Compton scattering.
  - a) What is the probability that an incident photon undergoes Compton scattering in the slab?
  - b) What is the probability that it undergoes photoelectric absorption?
  - c) What is the probability that an incident photon is not photoelectrically absorbed in the slab?
- 1.19 An atom of  $^{42}\text{K}$  (half-life = 12.4 h) has a probability of 0.894 of surviving 2 h. For a source that consists of five atoms,
  - a) what is the probability that all five will decay in 2 h and
  - b) what is the probability that none of the five atoms will decay in 2 h?
- 1.20 What are the answers to (a) and (b) of the last problem for a source of 100 atoms?



**Figure 1.2** Neutrons normally incident on a pair of slabs, 1 and 2. See Problems 1.21–1.24.

- 1.21 Monoenergetic neutrons are normally incident on a pair of slabs, arranged back to back, as shown in Figure 1.2. A neutron either is absorbed in a slab or else goes through without interacting. The probability that a neutron gets through slab 1 is  $1/3$ . If a neutron penetrates slab 1, then the probability that it gets through slab 2 is  $1/4$ . What is the probability that a neutron, incident on the pair of slabs, will
- traverse both slabs?
  - be absorbed in slab 1?
  - not be absorbed in slab 2?
- 1.22 If, in Figure 1.2, a neutron is normally incident from the right on slab 2, then what is the probability that it will
- be absorbed in slab 1?
  - not be absorbed in slab 2?
- 1.23 For the conditions of Problem 1.21, calculate the probability that a neutron, normally incident from the left, will
- not traverse both slabs,
  - not be absorbed in slab 1, and
  - be absorbed in slab 2.
- 1.24 What is the relationship among the three answers to the last problem and the corresponding answers to Problem 1.21?

